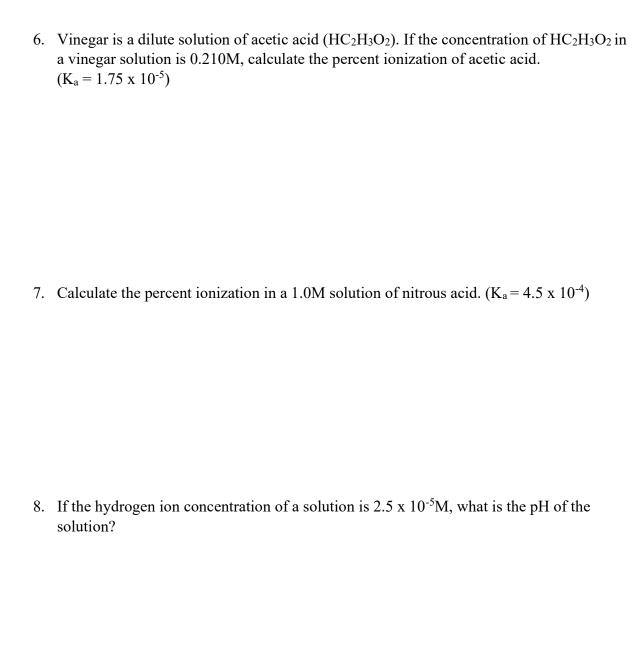
## Acid-Base Equilibria

1.	What is the formula for percent ionization?
2.	A solution of HCl has a pH of 2.3. Calculate the pOH, [H $^+$ ], and [OH $^-$ ] pOH = [H $^+$ ] = [OH $^-$ ] =
3.	What is the concentration of hydronium ions in a solution that has a volume of 2.50 L and 5.33 g of HCl? What is the pH of this solution?
4.	Calculate the pH of a solution prepared by mixing 0.15M HF with 0.18M NaF. ( $K_a$ of HF is 6.8 x $10^{-4}$ , p $K_a$ = 3.17)
5.	Calculate the pH of a solution that is $0.20M$ in HCHO <sub>2</sub> and $0.15M$ in NaCHO <sub>2</sub>



9. What is the pOH of a solution with a pH of 8.72?

10. What is the pH of a 0.50M ammonia, NH $_3$ solution? The $K_b$ of ammonia is 1.8 x 10 $^{\text{-5}}$ .