

Exam 1 Test Prep Vocab

Equilibrium	Saturated Solution	Dissolution	Rate	Activation Energy	Molecularity
Solute	Unsaturated Solution	Solubility	Le Chatelier's Principle	Van't Hoff Factor	Catalyst
Solvent	Rate Law	Substrate	Equilibrium Constant	Elementary Step	Reaction Mechanism
Intermolecular Forces	Colligative Properties	Transition State	Rate-Determining Step	Reversible Reaction	Intermediate

1. _____ - Each step in a chemical reaction
2. _____ - If a system at equilibrium is disturbed, it will react in such a way as to counteract the disturbance and return to equilibrium
3. _____ - Energy which a reacting species must have to form the transition state
4. _____ - A species that is formed and used up; will not appear in rate law
5. _____ - Number of molecules reacting in a step
6. _____ - The substance which undergoes reaction
7. _____ - The speed of a reaction
8. _____ - The component that is dissolved in a solution
9. _____ - (Activated complex) state corresponding to the highest energy along the reaction coordinate
10. _____ - Reactions that can proceed in either direction
11. _____ - Maximum amount of solvent dissolved at a certain temperature
12. _____ - Solution properties that depend upon only quantity of solute
13. _____ - The slowest step in a reaction
14. _____ - Ratio between the actual concentration of particles produced when the substance is dissolved and the concentration of a substance as calculated from its mass
15. _____ - a substance that speeds up the rate of the reaction without causing permanent change
16. _____ - Detailed sequence of reaction steps

- 17. _____ - Contains less than the optimal amount of solute
- 18. _____ - State when forward and reverse reactions become equal
- 19. _____ - The component of a solution that retains its original phase
- 20. _____ - The mathematical expression relating rate to concentration
- 21. _____ - The process of dissolving
- 22. _____ - Forces between a solvent and a solute
- 23. _____ - The maximum concentration that a solution can achieve with respect to a particular solute
- 24. _____ - K , the value of the reaction quotient at equilibrium