Balance the following redox reactions:

$$ClO_3^- + Cl^- \rightarrow Cl_2 + ClO_2$$

$$MnO_4^- + Fe^{2+} \rightarrow Mn^{2+} + Fe^{3+}$$

Calculate the standard cell potential:

$$2Ag^+ + Fe \rightarrow 2Ag + Fe^{2+}$$

$$Cu + 2Ag^+ \rightarrow Cu^{2+} + 2Ag$$

$$Zn + Fe^{2+} \rightarrow Fe + Zn^{2+}$$

Use the following reaction to answer the questions:
$Mn^{2+} + A1 \rightarrow Mn + A1^{3+}$
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Which of the above is the oxidizing agent:
Which of the above gets oxidized:
How many electrons are being transferred:
What is the cathode reaction:
Calculate the standard cell potential for the reaction:

Is the reaction spontaneous:	
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